

## **Cyclic olefin polymer as a novel membrane material for membrane distillation applications**

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## Abstract

A first attempt was made to prepare cyclic olefin polymer/copolymer (COP/COC) flat-sheet porous membranes by the well-known non-solvent induced phase separation method. In this study, two solvents (chloroform and 1,2,4-trichlorobenzene), different additives (polyvinylpyrrolidone, PVP, polyethylene glycol, PEG400, polyethylene oxide, PEO, and Sorbitan monooleate, Span 80) and coagulants (acetone and 70/30 wt% acetone/water mixture) were employed. The prepared membranes were characterized in terms of the thickness (70 - 85  $\mu\text{m}$ ), porosity ( $\sim 50 - 80\%$ ), liquid entry pressure (1.16 - 4.55 bar), water contact angle ( $\sim 86^\circ - 111^\circ$ ), mean pore size (158 – 265 nm), mechanical properties (tensile strength: 0.74 - 5.51 MPa, elongation at break: 3.34% - 7.94% and Young's modulus: 29 – 237 MPa), morphological and topographical characteristics. Short-term direct contact membrane distillation (DCMD) tests showed maximum permeate fluxes of 20  $\text{kg.m}^{-2}\text{h}^{-1}$  and 15  $\text{kg.m}^{-2}\text{h}^{-1}$  when using as feed distilled water and 30 g/L sodium chloride aqueous solution, respectively, with a high salt separation factor (99.99%). Long-term DCMD tests of some selected membranes carried out during 24 - 50 h showed that the membranes prepared with PEG additive exhibited more stable DCMD performance. In general, it was proved that COP can be successfully used as a novel polymer candidate in membrane distillation (MD) applications.

**Keywords:** Cyclic Olefin Polymer (COP); porous flat-sheet membrane; non-solvent induced phase separation; additive; direct contact membrane distillation.

## 1. Introduction

Water scarcity, one of the most critical issues of modern human life, has led many scientists to focus on water desalination and industrial wastewater reuse. Membrane-based technologies such as reverse osmosis (RO) and thermal desalination technologies such as multi effect distillation (MED) and multistage flash distillation (MSF) are currently being used. Compared to MED and MSF, RO is the preferred technology as it contributes with more than 60% in the desalination market. However, RO technology suffers some drawbacks such as its inability to treat high saline waters, brine discharge to the environment and high operating pressure requirement [1].

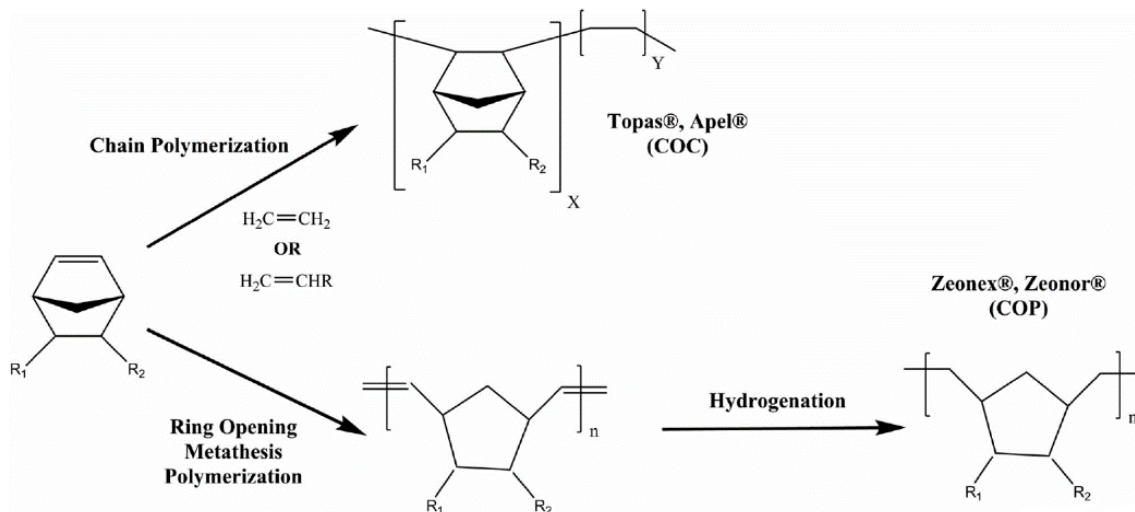
Membrane distillation (MD) is a non-isothermal separation technology, attractive in the field of desalination for its ability to treat high salinity water (up to 220 g/L NaCl) including RO brine up to its saturation [2]. Among other advantages, it exhibits ~100% rejection of non-volatile solutes and it can be easily integrated with renewable energy systems. Moreover, very low hydrostatic pressures, near atmospheric pressure, are applied. In MD technology, various configurations have been proposed. By applying either a temperature difference between both sides of a porous hydrophobic membrane (i.e. a hot feed and a cold permeate) or vacuum in the membrane permeate side, a transmembrane water vapor pressure difference is established giving rise to distinct MD variants (direct contact membrane distillation, DCMD; vacuum membrane distillation, VMD; air gap membrane distillation, AGMD; sweeping gas membrane distillation, SGMD, etc.). In all these configurations, only water vapor is transported through the membrane pores that must be maintained dry as the applied transmembrane hydrostatic pressure is lower than the water liquid entry pressure (*LEP*) of the membrane [3]. It is worth quoting that although the membrane has an important effect on the process efficiency, most MD studies have been focused on either theoretical modelling or the influence of different operating parameters on MD performance [4,5]. Only

during last decade, efforts have been done on MD membrane and module engineering [6]. Various innovative designs and novel membranes have been proposed for MD applications [7]. The MD membranes should exhibit high hydrophobicity or omniphobicity to reduce the risk of pore wetting by increasing the liquid entry pressure (*LEP*), high porosity to increase membrane permeability, optimized thickness and low thermal conductivity to enhance thermal efficiency, high chemical resistance and good thermal stability using appropriate materials to guarantee long-term membrane performance [8-10].

The hydrophobic polymers proposed for the preparation of MD membranes are polypropylene (PP) [11], polytetrafluoroethylene (PTFE) [12], PTFE-co-hexafluoropropylene (FEP) [13], polyethylene (PE) [14], polyvinylidene fluoride (PVDF) [15] and its derivatives (PVDF-co-tetrafluoroethylene (PVDF-TFE), PVDF-co-hexafluoropropylene (PVDF-HFP), PVDF-co-trifluorochloroethylene (PVDF-CTFE)) [16-18] which are fabricated by different processes. Some polymers such as PVDF have good processability by which the porous structure can be prepared via non-solvent induced phase separation (NIPS). However, others such as PP and PE are subjected to thermally induced phase separation (TIPS), and polymers such as PTFE and FEP need sintering and melt spinning, respectively [19].

Cyclic olefin polymers are relatively new types of olefin polymers and consequently show the characteristics of polyolefins. Cyclic olefin polymers exhibit interesting characteristics such as their high glass transition temperature ( $T_g$ : 80 - 180 °C) in combination with excellent transparency, good mechanical properties, low moisture absorption, low density, strength and hardness, little shrinkage, good electrical properties, low thermal conductivity (0.12 - 0.15 W.m<sup>-1</sup>. K<sup>-1</sup>), very good melt processability, biocompatibility and chemical resistance to acids, alkalis and polar organic solvents. These characteristics are attractive in different fields including packaging, optics and

medical tools. Cyclic olefin copolymer (COC), the copolymer of ethylene and norbornene, and cyclic olefin polymer (COP) are relatively new types of olefin polymers, although the first commercial product of COC appeared by the end of 1980s [20]. In 1984, Yamazaki et al. [21] from Zeon company in Japan developed a particular type of COP containing only norbornene, which has a rigid bridged-ring structure preventing crystallization. As can be seen in Fig.1, the structure of COP and COC are very similar [20].

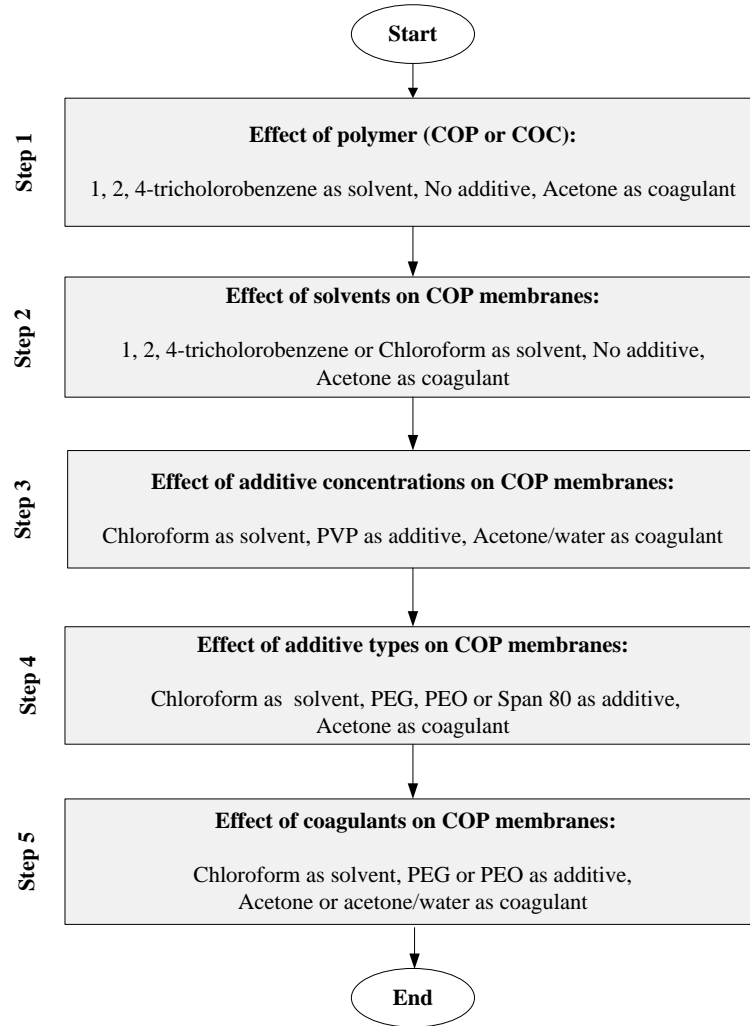


**Fig.1. Synthesis of cyclic olefin copolymers (COC) and cyclic olefin polymers (COP) from norbornene [20].**

It must be pointed out that there are very few reports on COC membranes and all of them deal with dense membranes [22,23]. For instance, gas transport, sorption properties and selectivity of a series of COC dense membranes prepared with different contents of norbornene (40-66%) were investigated for the first time by Hu et al. [22]. It was found that the norbornene content had a significant effect on the gas performance of dense COC membrane. Moreover, its fractional free volume increased with the norbornene content resulting in higher solubility and permeability. It was claimed that COC could be a suitable membrane material for the separation of  $O_2/N_2$  at

1 moderate pressures, less than 10 bar. Afterwards, Doğu and Ercan [23] evaluated the gas  
2 permeability and selectivity of COC composite membranes for H<sub>2</sub>/CO<sub>2</sub> and H<sub>2</sub>/CH<sub>4</sub> separations.  
3 These COC composite membranes were prepared by the melt processing method using different  
4 types of graphite nano-sheets. The O<sub>2</sub>/N<sub>2</sub> ideal selectivity was found to be higher than that of other  
5 traditional membrane materials and the permeability–selectivity tradeoff limit (Robeson's 2008  
6 upper bound) was surpassed for H<sub>2</sub>/CO<sub>2</sub> separation.

7 To the best of our knowledge, there is no documentation about COP and COC porous  
8 membranes. In this work, we propose for the first time the use of COP and COC to prepare porous  
9 membranes by NIPS method for different applications especially membrane distillation (MD). It  
10 is worth quoting that COP is a cheap and plastic polymer used for the first time as a membrane  
11 material. In this study, 5 steps were followed as shown in Fig.2 to investigate the effects of  
12 polymers, solvents, additive's concentration and type, and coagulants on the resultant membranes.  
13 The membranes were characterized by various techniques and the obtained results were compared  
14 with other membranes used in desalination by DCMD.



**Fig. 2. Schematic summary of the research study.**

## **2. Experimental**

### *2.1. Materials*

Cyclic olefin copolymer, Topas<sup>®</sup> 6013, and cyclic olefin polymer, Zeonex<sup>®</sup> 480R, were commercial grades and kindly donated by Topas Advanced Polymers GmbH and Zeon Europe GmbH (Germany), respectively. Their properties are shown in Table 1. Other used chemicals in this study include chloroform and 1,2,4-Trichlorobenzene as solvents, acetone as a non-solvent, Polyethylene glycol (PEG), Polyethylene oxide (PEO), Polyvinylpyrrolidone K90 (PVP) and,

Sorbitan monooleate (Span<sup>®</sup> 80), as additives. Isopropyl alcohol (IPA) and POREFIL<sup>®</sup> were employed as wetting liquids for the measurements of porosity and porometry, respectively. Sodium chloride (NaCl) and distilled water were used to prepare the salt aqueous feed solutions for DCMD experiments. The characteristics of these materials together with their corresponding suppliers are listed in Table 2.

**Table 1.** Properties of COP used in this study.

Commercial grade	Norbornene units (%)	$T_g^a$ (°C)	$\rho^a$ (g/cm <sup>3</sup> )	MFI <sup>a</sup> (g/10 min)
Topas <sup>®</sup> 6013	54.3	140	1.02	12 <sup>b</sup>
Zeonex <sup>®</sup> 480R	100	137	1.01	21 <sup>c</sup>

<sup>a</sup> MFI: melt flow index;  $T_g$ : glass transition temperature;  $\rho$ : density.

<sup>b</sup> at 260°C and 2.16 kg

<sup>c</sup> at 280°C and 2.16 kg

**Table 2.** List of used chemicals and solvents.

Material	Characteristics	Supplier
PEG	average $M_n$ : 400 g/mol	Sigma-Aldrich
PEO	$M_w$ : 200,000 - 300,000 g/mol	Sigma-Aldrich
PVP	$M_w$ : 360,000 g/mol	Fluka Chemie AG
Span <sup>®</sup> 80	non-ionic surfactant	Sigma-Aldrich
Chloroform	ACS reagent, puriss., $\geq 99.8\%$ , stabilized with ethanol	Acros Organics
1,2,4-Trichlorobenzene	Anhydrous, puriss., $\geq 99\%$	Sigma-Aldrich
Acetone	Puriss., $\geq 99.5\%$ , for analysis	Acros organics
POREFIL <sup>®</sup>	Chemical nature: pefluoroether Surface tension: 16 mN/m Vapor pressure: 3.33 Pa Viscosity: 4.4 mPa s	Porometer
IPA	Anhydrous, puriss., $\geq 99.5\%$	Sigma-Aldrich
NaCl	-	PanReac AppliChem GmbH



## 2.2. Membrane preparation

The solvent (chloroform or 1,2,4-trichlorobenzene) was mixed with the additive (0.1, 0.2 or 0.3 wt%) for 1h at room temperature (22°C) using a magnetic stirrer at 120 rpm. 10 wt% of polymers (COP, COC) was then added to the solution that was subsequently kept on the orbital shaker at 40°C and 100 rpm for almost 12 h until it became a homogeneous solution. Finally, the obtained polymer dope solution was degassed for 4 h at room temperature followed by casting onto a glass plate by using an automatic film applicator (Elcometer 4340, Elcometer®) with a knife gap of 0.25 mm and a velocity of 100 mm/s at room temperature (22°C). The cast film was immediately immersed in the non-solvent coagulation bath (acetone or 70/30 wt% acetone/water mixture) at room temperature. The flat-sheet membrane was taken out from the coagulants and rinsed with acetone/water mixture (70/30 wt%) to remove any residual solvent or additives. Finally, the membrane was air-dried for 24 h. Table 3 summarizes the different used solvents, coagulants, additives and their concentrations. The membrane code was based on the used polymers, solvents, additives, content of additive (wt%) and coagulants. For instance, the membrane PVP-0.2-AW means that it was prepared with chloroform as solvent, PVP as additive with 0.2 wt% and acetone/water as coagulants. All COP membranes were prepared with chloroform solvent except the membrane TCB-NA-A that was prepared with 1,2,4-Trichlorobenzene. The COC membrane was also prepared with TCB and named as COC-TCB-NA-A.

**Table 3.** Dope solution compositions and process parameters for the prepared membranes.

Membrane code	Solvent	Additive	Additive content (wt%)	Solvent content (wt%)	Coagulant
COC-TCB-NA-A	TCB <sup>a</sup>	-	-	90	Acetone
TCB-NA-A	TCB	-	-	90	Acetone
CF-NA-A	CF <sup>b</sup>	-	-	90	Acetone
PVP-0.1-AW	CF	PVP	0.1	89.9	Acetone/water <sup>c</sup>
PVP-0.2-AW	CF	PVP	0.2	89.8	Acetone/water
PVP-0.3-AW	CF	PVP	0.3	89.7	Acetone/water
PEG-0.2-A	CF	PEG	0.2	89.8	Acetone
PEG-0.2-AW	CF	PEG	0.2	89.8	Acetone/water
PEO-0.2-A	CF	PEO	0.2	89.8	Acetone
PEO-0.2-AW	CF	PEO	0.2	89.8	Acetone/water
Span 80-0.2-A	CF	Span 80	0.2	89.8	Acetone

<sup>a</sup> 1,2,4-trichlorobenzene, <sup>b</sup> chloroform, <sup>c</sup> 70/30 wt% in all experiments.

### 2.3. Membrane characterization

Different characterization techniques were applied to determine the characteristics of the prepared membranes and the results were compared to the general properties of the MD membranes.

The morphological structure of both the surface and cross-section of the membranes was investigated by the field emission scanning electron microscope (FE-SEM, MIRA3 TESCAN, Czech Republic) operating at 10 kV acceleration voltage. Previously, the cross-section of the samples was prepared by fracturing in liquid nitrogen. The membrane surface and the prepared

cross-section was sputter-coated by a thin gold layer using a sputter coater (Q150R S, Quorum, England)) during 60 s under 20 mA.

The membrane thickness was measured at 30 different spots on each sample using a micrometer equipped with a feeler (ISL Isocontrol). The average value of the measured thicknesses with their corresponding standard deviation was reported as the final thickness of each sample.

The hydrophobicity of the membrane surface was analyzed by water contact angle measurement. This was carried out at room temperature using a computerized optical system CAM100, equipped with a CCD camera, frame grabber and image analysis software CAM200usb. More information can be found elsewhere [24]. The volume of the drop was controlled by a Hamilton® glass syringe with a Hamilton® stainless steel needle, which was between 12 and 14  $\mu\text{L}$ . Five images were recorded during 4s for each drop and at least 10 drops were considered for each sample to determine the average water contact angle ( $\theta$ ) value together with its standard deviation.

The void volume fraction (i.e. porosity,  $\varepsilon$ ) of the flat-sheet membranes was determined by measuring the density of the membrane material ( $\rho_{pol}$ ) using isopropyl alcohol (IPA) and the density of the membrane ( $\rho_m$ ) using distilled water following the method described in [25]. According to this method, the weight of a pre-weighed sample was measured first in a water (dry membrane) and then in IPA (wetted membrane) by using a pycnometer. Finally, the porosity of sample was calculated based on the obtained density of the membrane material from the weight in IPA and the density of the membrane from the weight in water. The reported data for each membrane is the average of at least three measurements of different samples.

The membrane mean pore size was determined via wet/dry flow method using a gas-liquid displacement Porometer and its corresponding computer software (POROLUX™ 100, Porometer). Briefly, the flat-sheet membrane was first wetted by a wetting liquid (POREFIL®) and the gas

permeation flow (compressed air) was measured at room temperature for different transmembrane increasing hydrostatic pressures (0-0.7 MPa) to obtain the S-shaped wet curve. Subsequently, the dry curve was obtained by measuring the air flow rate under different transmembrane hydrostatic pressures through the dry sample. Both the cumulative filter flow (*CFF*) and the differential filter flow (*DFF*) curves were obtained using the above mentioned software. For each membrane, at least three tests were carried out. The followed procedure was described elsewhere [24,26].

The liquid entry pressure (*LEP*) of the membrane is the minimum hydrostatic pressure applied on the flat-sheet membrane before the liquid penetrates inside its pores. This was measured for distilled water according to the procedure and set-up described in [27]. These measurements were carried out using three different membrane samples and the average values together with their standard deviations were calculated.

The mechanical properties of the membranes were evaluated by measuring the tensile strength, elongation at break, and Young's modulus using a universal tensile tester (SANTAM STM20) equipped with a 6 N load cell, according to ASTM D 882. The tests were carried out at room temperature, a crosshead speed of 5 mm/min on rectangular membrane strips with dimensions of 50 mm × 10 mm. The thickness of each membrane sample was measured and considered in the calculation of the cross-sectional area to determine the tensile strength. For each membrane, five samples were considered and the average values together with their standard deviations were calculated.

Atomic force microscopy (AFM) was carried out to study the surface topographical structure and determine the roughness parameters (average roughness,  $R_a$ , and root mean square (RMS) roughness,  $R_q$ ) of both the top and bottom surfaces of the membranes using an Atomic Force Microscope (AFM, Ara research, model: Full plus).  $R_a$  is defined as the arithmetic average of the

absolute values of the surface height deviations measured from the mean plane and  $R_q$  is the root mean square average of the height deviations taken from the center plane. The measurements were conducted at room temperature over an area of  $5\text{ }\mu\text{m}\times 5\text{ }\mu\text{m}$ .

#### 2.4. Direct contact membrane distillation (DCMD)

The DCMD performance of the membranes were evaluated using two experimental set-ups, one for short-term experimental tests and the other for long-term experimental tests.

##### 2.4.1. Short-term experiments

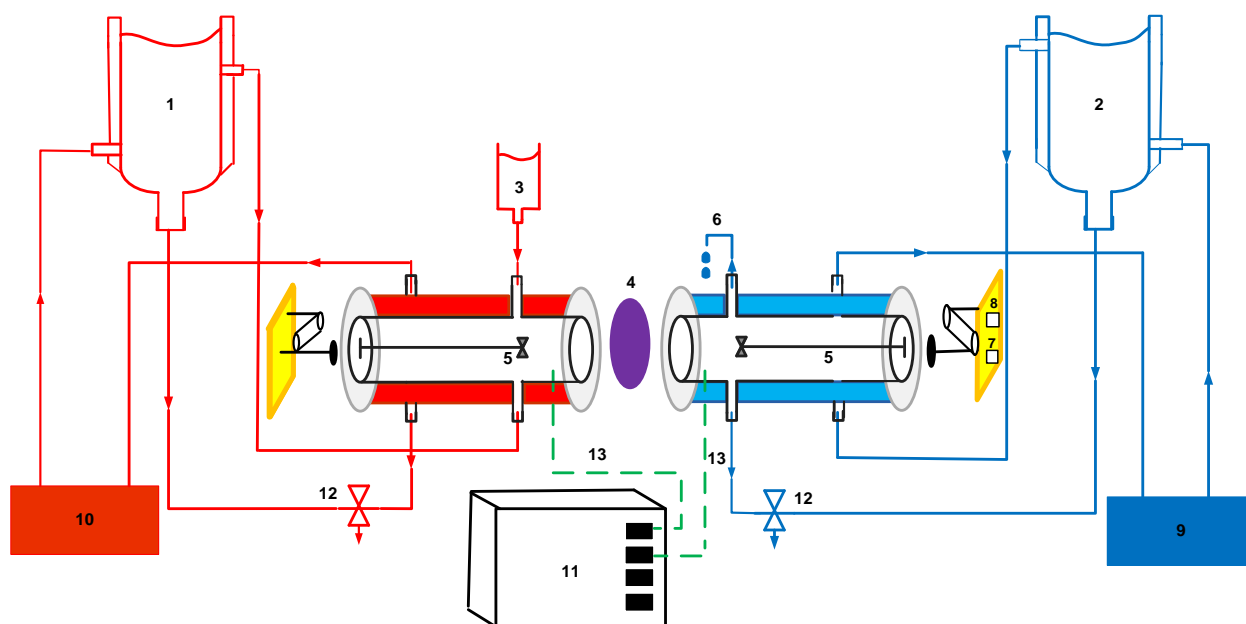
The short-term experiments of all COP membranes was carried out using the DCMD set-up schematized in Fig. 3. The permeate flux was measured using distilled water ( $J_w$ ) and 30 g/L NaCl aqueous solution ( $J$ ) as feed together with the salt separation factor ( $\alpha$ ). This system consists of two double-wall stainless steel cylindrical chambers (250 mL). The temperature inside each chamber was measured by temperature sensor (Pt-100) having an accuracy of  $\pm 0.1$  K. The feed and permeate were stirred at 700 rpm using a magnetic stirrer connected to a graduated motor. The membrane with the effective area of  $8.3\text{ cm}^2$  was placed between the two chambers. To measure the distilled water permeate flux ( $J_w$ ) both chambers were filled with distilled water. A checking leak test was carried out first by stirring distilled water in both the feed and permeate chambers at the same temperature (room temperature,  $22^\circ\text{C}$ ). This guarantees that no liquid water is transferred between the chambers and the hydrophobicity of the membrane is acceptable. A heating and a cooling system were connected to both chambers to control the feed and permeate temperatures at  $70^\circ\text{C}$  and  $25^\circ\text{C}$ , respectively. Then desalination test was performed. In this test, the feed chamber

1 was filled with salt solution and both the permeate flux ( $J$ ) and the salt separation factor ( $\alpha$ ) were  
2 calculated using the following equations:

$$J = \frac{\Delta m}{S \cdot \Delta t} \quad (1)$$

$$\alpha = 100 \times \left(1 - \frac{C_p}{C_f}\right) \quad (2)$$

3 where  $\Delta m$  is the registered mass of the produced water over time ( $\Delta t$ ),  $S$  is the surface area of the  
4 membrane, and  $C_p$  and  $C_f$  are the salt concentration of the permeate and feed, respectively, calculated  
5 from the previously calibration of the salt concentration *versus* the electrical conductivity. The electrical  
6 conductivity of the feed and permeate solutions were measured at the beginning and at the end of  
7 each DCMD test by a previously calibrated conductivity meter (Mettler-Toledo GmbH,  
8 SevenCompact S230). Two electrical conductivity sensors along with a temperature sensor  
9 InLab<sup>®</sup>731-ISM and InLab<sup>®</sup>741-ISM were used to analyze the salt content of the feed and  
10 permeate, respectively. With both sensors, a wide range of electrical conductivity (0.000  $\mu$ S/cm -  
11 1000 mS/cm) could be measured. The resolutions of the mentioned sensors for the permeate and  
12 feed are in the range of 0.001  $\mu$ S/cm and 0.01 mS/cm, respectively. The electrical conductivity of  
13 the permeate and feed at the beginning of each experiment was measured as  $0.972 \pm 0.001 \mu$ S/cm  
14 and  $48.96 \pm 0.01$  mS/cm, respectively.

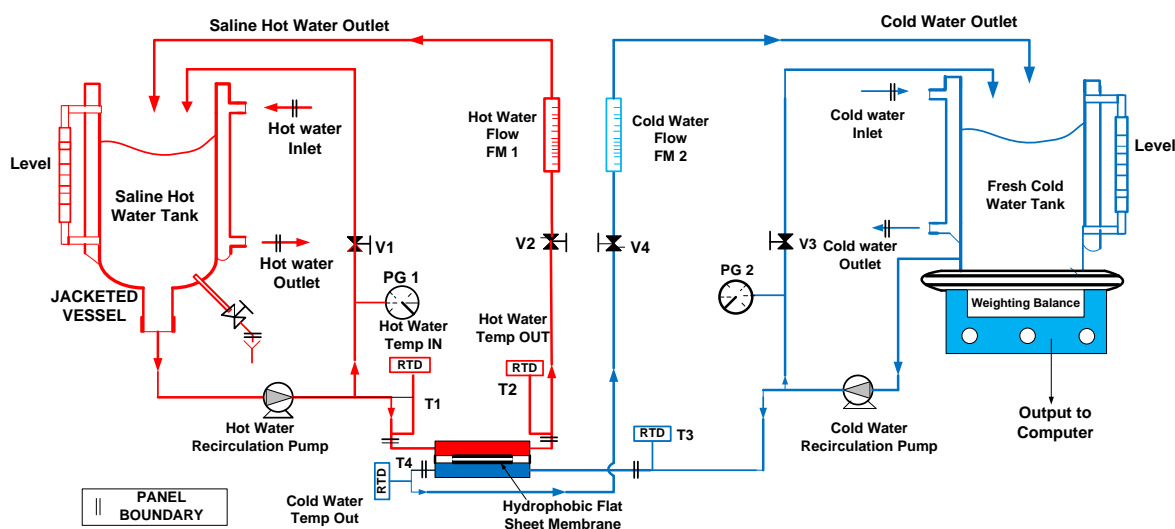


**Fig. 3.** DCMD short-term experiment set-up: (1) feed container, (2) permeate container, (3) feed supplier during DCMD test, (4) membrane holder, (5) stirrers, (6) pipette or tube for permeate flux measurement, (7) stirring rate regulator, (8) stirring graduated motor, (9) chiller, (10) thermostat, (11) temperature reader, (12) three way valves, (13) Pt-100 temperature sensors.

#### 2.4.2. Long-term experiments

Long-term experimental tests were performed for those membranes showing optimum DCMD performance during the previously taken short-term experimental tests. The DCMD set-up (Model No. FMD-SS-3090, M/s TECH INC, India) depicted in Fig. 4 was used. The effective membrane area of the rectangular module is  $27 \text{ cm}^2$ . The DCMD experiment was carried out using 30 g/L NaCl aqueous solution as feed and distilled water as permeate. The volume of the feed and permeate tanks is 5 L. Both the feed and permeate circulated at a flow rate of 30 L/h using two circulation pumps (TECH INC, India). Similar to the short-term experimental test, both the feed and permeate tanks were connected to the heating and cooling systems to adjust the feed and

1 permeate temperatures to 70 and 25°C, respectively. During the DCMD test, the mass of the  
 2 produced water in the permeate tank was weighed every 60s using a digital balance  
 3 (ACCUMAX™) connected to a computer, which also permitted to record the inlet and outlet  
 4 temperatures of the membrane module. The separation factor was evaluated from the  
 5 measurements of the electrical conductivity of the feed and permeate solutions as stated  
 6 previously.



**Fig. 4.** Schematic and real images of the DCMD long-term experimental set-up.



### 3. Results and discussion

#### 3.1. Study of the affinity between polymer, solvent and non-solvent

The correct selection of a solvent for a given polymer is necessary to obtain a homogenous polymer solution. The identification of a suitable solvent for a polymer can be facilitated by the knowledge of their solubility parameters. Hansen solubility parameter (HSP) is typically used for selecting proper solvents for polymers. This parameter ( $\delta$ ) is characterized by three intermolecular interactions, hydrogen bonding interaction ( $\delta_h$ ), polar interaction ( $\delta_p$ ) and nonpolar (dispersive) interaction ( $\delta_d$ ) as shown by the following relation [28]:

$$\delta^2 = \delta_h^2 + \delta_p^2 + \delta_d^2 \quad (3)$$

A suitable solvent for a given polymer should have a solubility parameter close to that of the polymer. Therefore, the Hansen solubility parameter distance ( $R_{HSP}$ ) is another measure to determine a good solvent ( $S$ ) for a given polymer ( $P$ ).  $R_{HSP}$  can be calculated as follows:

$$R_{HSP}^2(P-S) = 4(\delta_{dP} - \delta_{dS})^2 + (\delta_{pP} - \delta_{pS})^2 + (\delta_{hP} - \delta_{hS})^2 \quad (4)$$

More affinity between two chemicals is identified by a low  $R_{HSP}$  value between them.

The HSP of mixed non-solvents can be determined as follows [29]:

$$[\delta_{d_{NS}}, \delta_{p_{NS}}, \delta_{h_{NS}}] = [a\delta_{d1} + b\delta_{d2}, a\delta_{p1} + b\delta_{p2}, a\delta_{h1} + b\delta_{h2}] / (a+b) \quad (5)$$

where the subscripts 1 and 2 are for each compound present in the non-solvent mixture (i.e. acetone and water) and  $a$  and  $b$  are the volume fractions of different components in the non-solvent mixture. Therefore, to select a coagulant for a given polymer solution, the  $R_{HSP}$  (NS-S) parameter was considered as:

$$R_{HSP}^2(NS-S) = 4(\delta_{d_{NS}} - \delta_{d_S})^2 + (\delta_{p_{NS}} - \delta_{p_S})^2 + (\delta_{h_{NS}} - \delta_{h_S})^2 \quad (6)$$

The lower is the  $R_{HSP}$  (NS-S) value, the weaker is the coagulant for a polymer solution.

The HSP components of all used pure materials are listed in Table 4 together with the calculated  $R_{HSP}$  values [28,30]. In the calculations of  $R_{HSP}$  of P-S and NS-S, the HSP of the additive was ignored because of its very small amount in the polymer solution and the HSP of COP was assumed to be close to COC (Topas<sup>®</sup> 6013) [31]. Based on the obtained  $R_{HSP}$  values, since the  $R_{HSP}$  of P-CF (3.7) is lower than the  $R_{HSP}$  of P-TCB (5.4), the polymer has more affinity to chloroform than to 1,2,4- trichlorobenzene. Also, acetone-chloroform mixture ( $R_{HSP}$  of acetone-CF=8.7) is more compatible compared to acetone/water-chloroform mixture ( $R_{HSP}$  of Acetone/water-CF=11.1) as shown in Table 4. So, it is expected that the exchange rate of solvent and non-solvent in the acetone-chloroform mixture is faster than acetone/water-chloroform mixture.

**Table 4.** Solubility parameters of the used materials in this study with their corresponding  $R_{HSP}$  values.

Material type	Solubility parameters (MPa <sup>1/2</sup> )				$R_{HSP}$				
	$\delta_d$	$\delta_p$	$\delta_h$	$\delta$	P-S	NS-S	A-S	A-NS	A-NS <sub>m</sub>
Polymer (COP/COC)	18.0	3.0	2.0	18.3	-	-			
1,2,4-Trichlorobenzene	20.2	6.0	3.2	21.3	5.4	11.1 <sup>a</sup>			
Chloroform	17.8	3.1	5.7	18.9	3.7	8.7 <sup>a</sup>			
Acetone	15.5	10.4	7.0	19.9	-	-			
Water	15.5	16.0	42.3	47.8	-	-			
Acetone/water (70/30)	15.5	12.1	17.6	26.4	-	15.6 <sup>b</sup>			
PVP	17.4	8.8	14.9	24.5			10.85 <sup>b</sup>	-	5.71
PEG400	19.18	3.49	3.58	19.8			3.5 <sup>b</sup>	10.65 <sup>a</sup>	18.02
PEO	17.8	0.56	9.1	19.9			4.1 <sup>b</sup>	11.06 <sup>a</sup>	15.05
Span 80	17.2	7.7	10.3	21.5			6.6 <sup>b</sup>	5.5 <sup>a</sup>	-

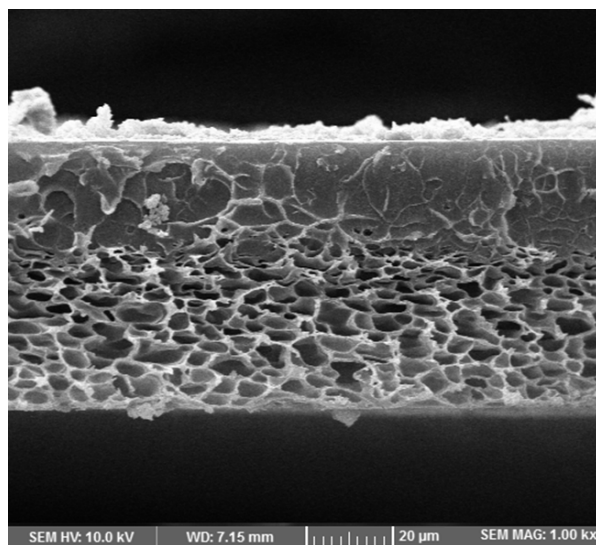
P: polymer, S: solvent, NS: non-solvent, A: additive; NS<sub>m</sub>: mixed non-solvent (acetone/water)

<sup>a</sup> Acetone as non-solvent (NS)

<sup>b</sup> Chloroform as solvent (S)

### 3.2. Influence of the polymer type (COC and COP) on membrane structure and performance

As mentioned before, preliminary studies were carried out using both COC and COP polymers in order to select the adequate polymer to prepare porous membranes. COC is a copolymer of ethylene and norbornene whereas COP is a polymer composed of only norbornene monomer. First, COC was dissolved in TCB as it is explained previously and the resultant membrane (COC-TCB-NA-A) was tested in DCMD short-term set-up using distilled water as feed at 70°C maintaining the permeate temperature at 25 °C. No permeate flux was detected. The FESEM image of the cross-section of this membrane (Fig. 5) showed that a thick dense layer was formed on the top of the membrane surface. This may be related to polyethylene, which is present in the backbone chain. So, TIPS method is required for the preparation of a completely porous structure from COC similar to the preparation of porous polyethylene membrane. Based on this observation, COP was selected for further studies as it is shown later on.



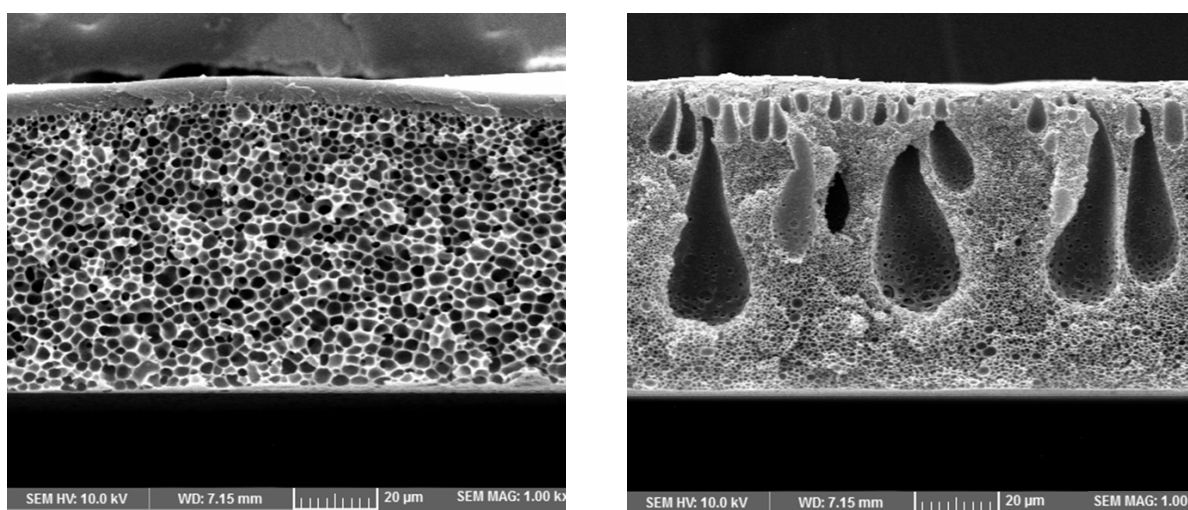
**Fig. 5.** FESEM image of the cross section of COC-TCB-NA-A membrane.

### 3.3. Solvent effect on COP membrane structure and performance

The COP membranes (TCB-NA-A and CF-NA-A, Table 3) were prepared using two different solvents, CF, which has more affinity to COP and TCB, which exhibits less affinity to COP (Table 4). Acetone was used as a non-solvent to prepare both membranes. No DCMD permeate flux could be detected for TCB-NA-A when using distilled water as feed and a bulk temperature difference of 45°C. However, CF-NA-A shows a permeate flux of 20 kg.m<sup>-2</sup>.h<sup>-1</sup> when distilled water was used as feed (Table 7). According to the cross section FESEM images of TCB-NA-A and CF-NA-A presented in Fig. 6, the structure of TCB-NA-A is uniformly spongy consisting of many small closed cells and a top dense skin layer, whereas the membrane CF-NA-A exhibits the characteristics of an asymmetric phase inversion membrane structure, consisting of a thick finger-like structure, a top denser skin layer and a sponge-like bottom layer. The presence of the top dense layer of the membrane TCB-NA-A prevented water production by DCMD. It is worth noting that  $R_{SHP}$  related to NS-S interaction for TCB-NA-A is higher than that of CF-NA-A. The low affinity between the non-solvent and solvent for TCB-NA-A resulted in slower inter-diffusion between them, a delayed COP coagulation and a sponge-like structure as a consequence. In general, during polymer solidification, the composition near the top surface of the membrane shifts to the polymer-rich phase resulting in a denser top layer than that of the bottom. The opposite occurred for CF-NA-A membrane resulting in a finger-like structure formation. The effect of solvent evaporation on membrane structure was investigated by Yamasaki et al. who claimed that more solvent evaporation time resulted in a thicker skin layer and more uniform structure [32]. Moreover, Friedrich et al. [33] investigated the effect of different solvents viscosity on the membrane structure and reported that the lower the solution viscosity resulted in a higher occurrence of cavities through the membrane. Therefore, in comparison to TCB, CF having lower viscosity and

more affinity to polymer for better solubility and to coagulant (i.e. acetone) for higher exchange rate of solvent and non-solvent was selected as solvent for next studies.

Both COC and COP are polyolefins. As mentioned in the previous section, COC containing polyethylene in its backbone chain requires TIPS method for the preparation of a completely porous structure while COP acquires the characteristics of polyolefin and therefore it is possible to produce COP porous structure by NIPS method (CF or TCB-NA-A).

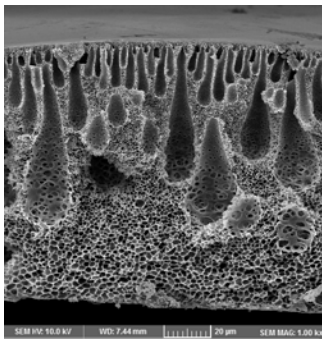
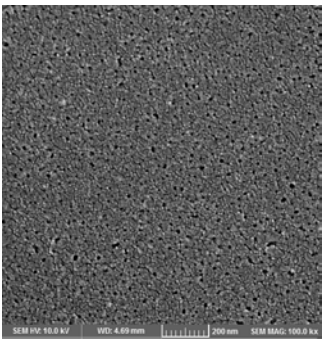
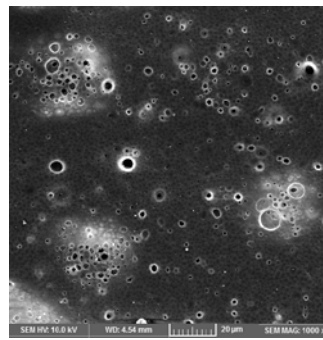
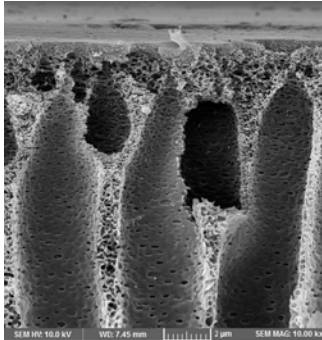
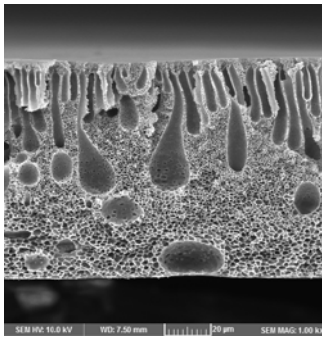
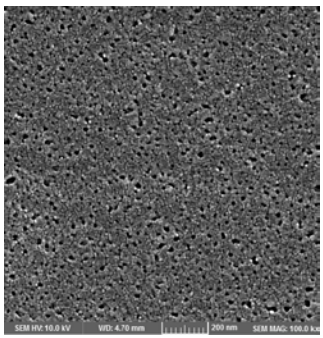
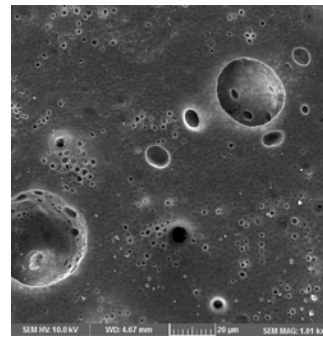
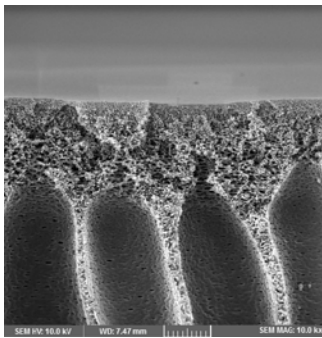
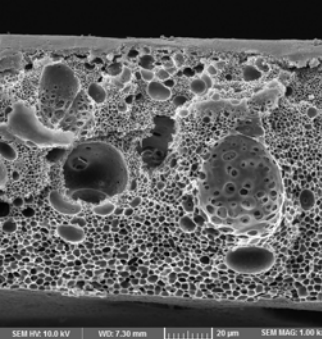
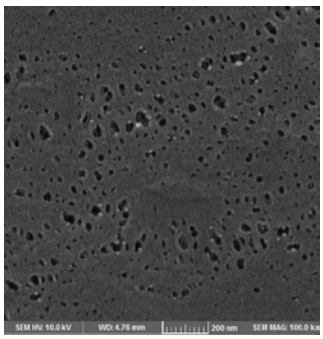
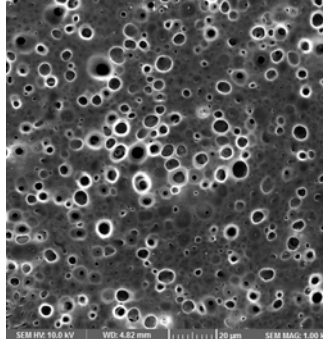
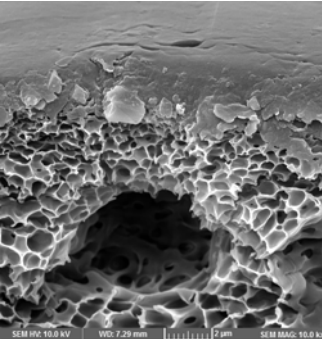


**Fig. 6.** Cross section FESEM image of TCB-NA-A (left) and CF-NA-A (right) membranes.

#### *3.4. Effect of PVP concentration on COP membrane structure and DCMD performance*

To study the effect of the additive concentration on the membrane structure, three COP membranes (PVP-0.1, 0.2, 0.3-AW) were prepared with different amounts of PVP (0.1, 0.2 and 0.3 wt%). Being a hydrophilic polymer, PVP is a well-known pore former used to prepare polymeric phase inversion porous membranes. Fig. 7 shows the FESEM images of the top, bottom and cross section of the PVP- 0.1, 0.2, 0.3-AW membranes. The three membranes exhibit a porous top layer, a finger-like structure (PVP-0.1-AW, PVP-0.2-AW) or macrovoids (PVP-0.3-AW) cross-section and a spongy bottom layer. These observations can be explained by means of the

concept of phase separation rate, which is influenced by the complex correlation between the thermodynamic enhancement and rheological hindrance inducing or suppressing the formation of macrovoids or finger-like structure [34]. The addition of PVP into the polymer solution brought an enhancement of demixing as a result of thermodynamic instability resulting in void or finger-like structure formation. On the other hand, the effect of PVP on COP membrane structure may be also explained using Hansen solubility parameter. The  $R_{HSP}$  values for different additives-solvent/non-solvents were determined and the obtained results were also listed in Table 4. According to the obtained data, the affinity of PVP-acetone/water mixture is more than the affinity of PVP-CF and consequently resulted in the rapid outflow of PVP and caused finger like structure. However, the increase of PVP content in the casting solution (PVP-0.3-AW) caused an enhancement of the viscosity and reduced the rate of solvent-non solvent exchange or which is the same a delay of demixing. In this case, the predominant kinetic effect of the viscosity enhancement might contribute to the suppression of the finger-like structure, an enhancement of the top skin layer and the formation of macrovoids between the top and the spongy bottom layer. This was also observed by Jung et al. [35] who found that the top membrane skin layer became thicker and the finger like macrovoids disappeared with the increase of PVP concentration in the polymer solution and this effect was more obvious in the case of high molecular weight of PVP because of its low solubility in water. It seems that more amount of high molecular weight of PVP (i.e. PVP-0.3-AW) caused macrovoids rather than PVP-0.1-AW and PVP-0.2-AW. It may also have suggested that more PVP are located at PVP-0.3-AW membrane and could take a room in the pores [35].

Membrane Code	Cross section ( $\times 1000$ magnification)	Top surface ( $\times 100,000$ magnification)	Bottom surface ( $\times 1000$ magnification)	Cross section near top surface ( $\times 10,000$ magnification)
PVP-0.1-AW				
PVP-0.2-AW				
PVP-0.3-AW				

1 **Fig. 7.** FESEM images (cross section, top and bottom surfaces, and details of the cross section near the top surface) of the membranes prepared  
2 with different PVP concentrations (0.1-0.3 wt%): PVP-0.1-AW, PVP-0.2-AW and PVP-0.3-AW.

Table 5 summarizes the structural properties of all COP membranes prepared using chloroform as solvent in this study. The COP membrane without any additive (CF-NA-A) as well as the COP membranes prepared with PVP as additive showed a good porosity with values greater than 74.8%. It is worth quoting that the high rate of demixing is, the greater porosity is achieved [36]. El-Bourawi et al. [8] reported that the porosity of MD membranes lies in the range of 30-85%. Therefore, from the porosity point of view, all prepared membranes in this study are suitable for MD membranes.

The thickness, *LEP*, contact angle and mean pore size of COP membranes are also listed in Table 5. No significant difference could be drawn from the thickness of the membranes if the corresponding standard deviations are taken into consideration. The porosity of the membranes increased while the *LEP* and contact angle decreased when a larger amount of PVP was used as additive (PVP-0.1, 0.2, 0.3-AW in Table 5). The mean pore size was increased with the increase of the PVP concentration followed by a clear reduction. This decrease of the pore size may be due to the increase of the dense skin layer as it can be seen from the top surface of PVP-0.3-AW membrane (Fig. 7).

The membrane CF-NA-A prepared without any additive exhibits the highest water contact angle, 111°, showing the good hydrophobic character of COP material. All membranes, except PVP-0.3-AW, have a water contact angle greater than 90° confirming their hydrophobicity. As it was expected, the addition of the hydrophilic additive PVP resulted in a decrease of the water contact angle (see CF-NA-A and PVP-0.1, 0.2, 0.3-AW membranes in Table 5).

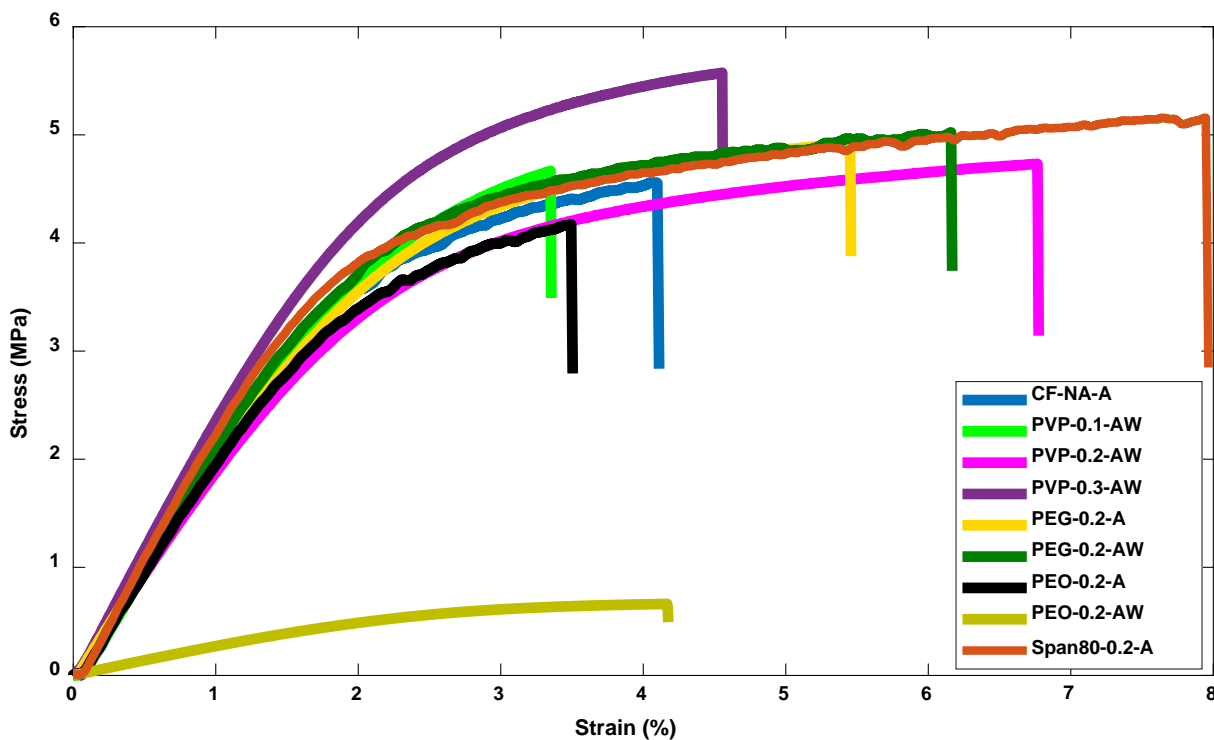
Although MD process is carried out under hydrostatic pressures near atmospheric pressure to guarantee an adequate membrane packing in modules and their subsequent stable long-term operation, the membrane must exhibit appropriate mechanical properties. The obtained stress-strain curves of the prepared COP membranes were plotted in Fig. 8. The tensile strength ( $\sigma_b$ ), elongation at break ( $\varepsilon_b$ ) and Young's modulus (E) were calculated and the results were given in Table 5.



1 **Table 5.** Thickness, porosity ( $\varepsilon$ ),  $LEP$ , water contact angle ( $\theta$ ), mean pore size and mechanical properties (tensile strength,  $\sigma_b$ ,  
2 elongation at break,  $\varepsilon_b$ , and Young's modulus,  $E$ ) of the prepared COP membranes.

Membrane code	Thickness ( $\mu\text{m}$ )	$\varepsilon$ (%)	$LEP$ (bar)	$\theta$ ( $^\circ$ )	Mean pore size (nm)	$\sigma_b$ (MPa)	$\varepsilon_b$ (%)	$E$ (MPa)
<b>CF-NA-A</b>	$85 \pm 5$	$80.3 \pm 3.3$	$1.96 \pm 0.19$	$111.0 \pm 3.2$	$211 \pm 3$	$4.53 \pm 0.02$	$4.1 \pm 1.4$	$208 \pm 6$
<b>PVP-0.1-AW</b>	$75 \pm 5$	$74.8 \pm 1.5$	$2.65 \pm 0.12$	$92.1 \pm 1.7$	$201 \pm 5$	$4.67 \pm 0.36$	$3.3 \pm 1.7$	$215 \pm 5$
<b>PVP-0.2-AW</b>	$70 \pm 5$	$79.4 \pm 1.5$	$1.67 \pm 0.18$	$93.5 \pm 2.6$	$265 \pm 6$	$4.73 \pm 0.32$	$6.8 \pm 1.7$	$190 \pm 4$
<b>PVP-0.3-AW</b>	$75 \pm 5$	$82.3 \pm 1.3$	$1.16 \pm 0.32$	$86.3 \pm 1.5$	$205 \pm 2$	$5.51 \pm 0.39$	$4.5 \pm 1.2$	$237 \pm 9$
<b>PEG-0.2-A</b>	$75 \pm 5$	$75.8 \pm 1.8$	$2.82 \pm 0.09$	$102.9 \pm 2.4$	$196 \pm 8$	$4.82 \pm 0.19$	$5.4 \pm 1.8$	$210 \pm 7$
<b>PEG-0.2-AW</b>	$75 \pm 5$	$73.3 \pm 2.2$	$2.04 \pm 0.14$	$103.8 \pm 1.4$	$158 \pm 5$	$5.05 \pm 0.04$	$6.1 \pm 1.8$	$222 \pm 6$
<b>PEO-0.2-A</b>	$80 \pm 5$	$69.5 \pm 2.9$	$3.10 \pm 0.26$	$98.8 \pm 1.2$	$254 \pm 12$	$4.19 \pm 0.03$	$3.5 \pm 0.4$	$204 \pm 6$
<b>PEO-0.2-AW</b>	$75 \pm 5$	$61.3 \pm 1.6$	$2.75 \pm 0.19$	$99.7 \pm 2.7$	$241 \pm 9$	$0.74 \pm 0.04$	$4.2 \pm 1.7$	$29 \pm 1$
<b>Span 80-0.2-A</b>	$80 \pm 5$	$50.0 \pm 1.7$	$4.55 \pm 0.34$	$100.0 \pm 2.2$	$239 \pm 2$	$5.15 \pm 0.29$	$7.9 \pm 1.0$	$232 \pm 5$

As can be seen in Fig. 8, all samples showed an elastic deformation at the beginning of the stress-strain curve followed by a plastic deformation. By increasing the amount of PVP, the tensile strength of the membrane gradually increased up to 5.51 MPa for the membrane PVP-0.3-AW while the elongation at break reached the optimum value of 6.76% for the membrane PVP-0.2-AW and then decreased. This reduction may be attributed to the increase of the porosity of the membrane PVP-0.3-AW and the formed macrovoids through its cross-section (Fig. 7). Also, the Young's modulus for PVP modified membrane represents the minimum value of 190 MPa for the membrane PVP-0.2-AW, due to its large finger-like structure. Moreover, the thick skin layer of the membrane PVP-0.3-AW may be responsible for its maximum E value. Simone et al. [37] investigated the effect of different amounts of PVP on the mechanical properties of hollow fiber PVDF membrane. They reported that by increasing the PVP content, the Young's modulus increased and then dropped off and further increment of PVP improved the mechanical strength. Alayande et al. [38] observed a higher porosity and a lower tensile strength with the addition of PVP to the dope solution. In general, compared to other MD membranes, all prepared COP membranes showed reasonably good mechanical properties. Tomaszewska [39] reported a strength at break of 1.6 MPa for a PVDF membrane that was decreased with the addition of LiCl due to macrovoid formation. Hou et al. also reported the stress at break for hollow fiber PVDF membrane of 4.66 MPa and elastic modulus of 160 MPa [40]. However, compared to these membranes, the prepared COP membranes exhibited less elongation at break. This is related mainly to the plastic behavior of COP and its application under  $T_g$  resulting in more brittle membranes.



**Fig. 8.** Stress-strain curves of the prepared COP membranes.

Fig. 9 and Table 6 show the 3D AFM images and roughness parameters ( $R_a$  and  $R_q$ ) of both the top and bottom surfaces of all COP membranes prepared with chloroform as solvent. All AFM images exhibited smooth nodule-like and valley-like structures. When the content of the PVP was increased (PVP-0.1, 0.2, 0.3-AW) inhomogeneous roughness was observed with some deep valleys and high peaks in some parts of the top membrane surface. The determined roughness parameters were higher for the membranes prepared with greater PVP amount. With the increase of PVP from 0.1 wt% to 0.2 wt%, the roughness parameters and contact angles were maintained almost the same taking into consideration the corresponding standard deviations. It seems that the roughness increase counteracts the hydrophilic characteristic of PVP. A significant increase of  $R_a$  and  $R_q$  values was observed when the PVP content was increased from 0.2 wt% to 0.3 wt% for both the top (47% and 44%, respectively) and bottom (56% and 66%, respectively) surfaces. However,  $\theta$  values were decreased from 93.5° to 86.3° due to the hydrophilic character of PVP,

which overcomes the effect of the roughness increase. It is worth mentioning that PVP-modified membranes contain smooth surface along with intensified asperities in some parts of the membrane surface (Fig. 9). In addition, the roughness parameters of the bottom surface are greater than the top surface. This is due to the sponge-like structure of the bottom surface caused by the slower phase inversion near the polymer/glass interface compared to the air/polymer interface [41].

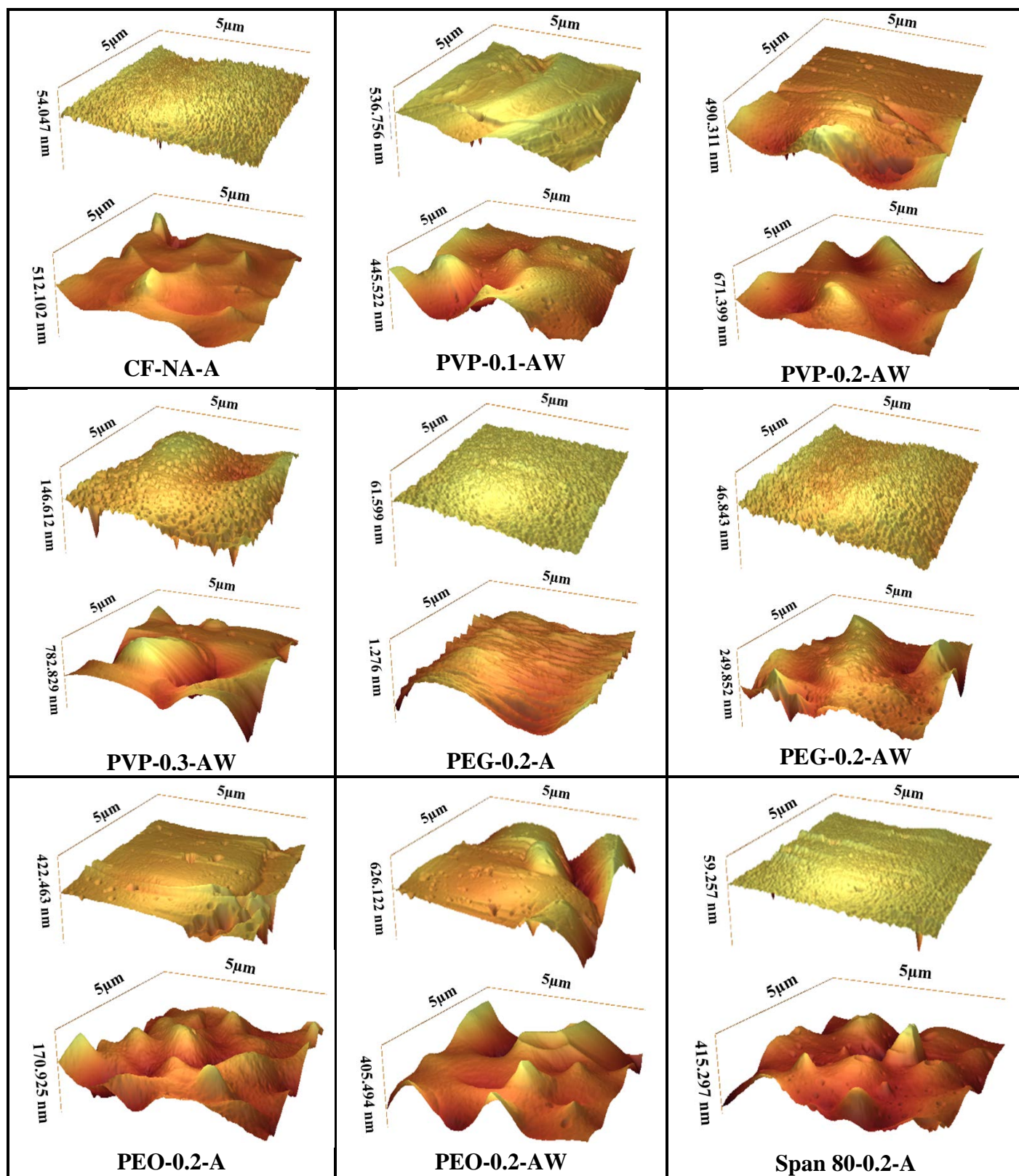
**Table 6.** Roughness parameters ( $R_a$ ,  $R_q$ ) together with the corresponding standard deviations of the top and bottom surfaces of the prepared COP membranes (scan range considered  $5\ \mu\text{m} \times 5\ \mu\text{m}$ ).

Membrane code	Top surface		Bottom surface	
	$R_a$ (nm)	$R_q$ (nm)	$R_a$ (nm)	$R_q$ (nm)
<b>CF-NA-A</b>	$106.4 \pm 3.8$	$126.7 \pm 5.0$	$108.1 \pm 0.5$	$132.1 \pm 4.0$
<b>PVP-0.1-AW</b>	$99.6 \pm 6.0$	$129.2 \pm 18.1$	$225.4 \pm 2.1$	$265.2 \pm 7.3$
<b>PVP-0.2-AW</b>	$105.9 \pm 10.0$	$131.4 \pm 18.9$	$237.5 \pm 7.2$	$258.4 \pm 5.2$
<b>PVP-0.3-AW</b>	$155.6 \pm 12.0$	$189.6 \pm 18.2$	$369.7 \pm 20.1$	$428.2 \pm 18.5$
<b>PEG-0.2-A</b>	$81.4 \pm 2.0$	$94.7 \pm 2.5$	$237.2 \pm 2.4$	$278.0 \pm 5.3$
<b>PEG-0.2-AW</b>	$99.8 \pm 14.0$	$116.7 \pm 12.7$	$248.8 \pm 11.2$	$298.3 \pm 17.5$
<b>PEO-0.2-A</b>	$99.0 \pm 16.5$	$121.9 \pm 15.8$	$212.5 \pm 5.6$	$253.9 \pm 18.3$
<b>PEO-0.2-AW</b>	$131.6 \pm 2.0$	$159.2 \pm 2.2$	$236.3 \pm 6.4$	$290.1 \pm 15.6$
<b>Span 80-0.2-A</b>	$107.9 \pm 2.8$	$129.3 \pm 4.8$	$120.3 \pm 3.6$	$152.3 \pm 6.5$

The obtained short-term DCMD results are shown in Table 7. No clear trend could be plotted between the permeate flux and the PVP concentration in the COP solution. With the increase of the PVP concentration from 0.1 to 0.2 wt%, the permeate flux was increased and with a further increase of PVP concentration up to 0.3 wt% it was decreased. This can be related with the pore

size as it was discussed before (Table 5). Moreover, the surface roughness affects the MD performance so that the higher surface roughness, the lower temperature and concentration polarization (TCP). The enhancement of the surface roughness might contribute to the reduction of the TCP effect attributed to the generated micro-turbulences at the membrane surface and might increase the DCMD permeate flux. Compared to the membrane PVP-0.1-AW, some improvement in roughness parameters of the membrane PVP-0.2-AW resulted in a little enhancement of the water contact angle. Therefore, this effect together with the increase of the mean pore size resulted in some improvements of the permeate flux (Table 7). However, the high  $R_a$  value of the membrane PVP-0.3-AW was not related to the DCMD permeate flux improvement (Table 7) because the mean pore size dropped from 265 nm to 205 nm (see Table 5). The slight decline of the salt separation factor observed for the PVP-0.3-AW membrane can be due to the reduction of the *LEP* (Table 5). All PVP-0.3-AW membrane results seem to indicate that an increase of the PVP content from 0.2 to 0.3 wt.% reduced the MD performance because bigger macrovoids appeared through its cross-section (see Fig. 7) and these are well-known by their reduction of the permeate flux.

It is worth noting that the highest permeate flux in short-term DCMD was obtained for the membrane CF-NA-A with a high salt separation factor (99.99%). However, it is essential to track its performance in long-term DCMD, too. Among all prepared COP membranes with additives, PVP-0.2-AW membrane exhibited the highest short-term DCMD performance. Based on these DCMD results, in order to compare the effect of different types of additives on the COP membrane structure and DCMD performance, the amount of all additives was fixed at 0.2 wt%.



1 **Fig. 9.** 3D AFM images of the top (up) and bottom (down) surfaces of the prepared COP  
2 membranes.

1 **Table 7.** Short-term DCMD performance.

Membrane code	Distilled water flux ( $J_w$ ) (kg/m <sup>2</sup> h)	Permeate flux ( $J$ ) (kg/m <sup>2</sup> h)	Electrical conductivity of permeate (μS/cm)			Electrical conductivity of feed (mS/cm)			Salt separation factor ( $\alpha$ ) (%)
			Test 1	Test 2	Test 3	Test 1	Test 2	Test 3	
<b>CF-NA-A</b>	20.0±0.5	15.0±1.0	4.726±0.002	5.103±0.001	5.241±0.001	46.63±0.02	46.55±0.03	46.72±0.02	99.9942±0.0004
<b>PVP-0.1-AW</b>	11.0±0.7	9.0±0.9	4.361±0.001	3.988±0.002	4.232±0.001	46.78±0.02	47.08±0.03	46.98±0.04	99.9955±0.0006
<b>PVP-0.2-AW</b>	16.0±1.0	13.0±1.3	5.553±0.002	5.932±0.001	6.141±0.003	46.54±0.01	46.49±0.02	46.51±0.02	99.9928±0.0005
<b>PVP-0.3-AW</b>	8.4±1.2	7.4±1.4	10.92±0.002	11.68±0.002	12.011±0.001	46.50±0.01	46.35±0.02	46.43±0.01	99.9837±0.0006
<b>PEG-0.2-A</b>	14.0±1.0	12.3±1.1	2.796±0.002	3.315±0.002	3.535±0.002	47.45±0.02	47.13±0.02	47.23±0.02	99.9972±0.0006
<b>PEG-0.2-AW</b>	11.0±0.8	8.0±1.0	9.578±0.001	10.257±0.001	10.860±0.002	46.51±0.02	46.36±0.02	46.47±0.01	99.9858±0.0010
<b>PEO-0.2-A</b>	9.0±1.1	7.3±0.9	3.411±0.004	3.107±0.003	3.732±0.002	46.92±0.02	47.17±0.01	46.89±0.01	99.9968±0.0005
<b>PEO-0.2-AW</b>	3.1±0.8	2.2±0.6	13.562±0.001	12.841±0.002	14.215±0.001	46.39±0.02	46.41±0.02	46.30±0.01	99.9805±0.0011
<b>Span 80-0.2-A</b>	5.3±1.6	3.0±0.8	2.882±0.001	2.109±0.002	1.899±0.001	46.88±0.01	46.68±0.02	47.12±0.01	99.9986±0.0008

2

### 3.5. Study of different types of additives and coagulants

The characteristics and DCMD performance of the prepared porous COP membranes were also summarized in Tables 5 and 7, respectively. First, the effects of the additives on the membrane morphology, structural properties and DCMD performance are discussed, then the effects of the coagulants on the same membrane properties.

#### 3.5.1 Effects of different additives on membrane characteristics and short-term DCMD performance

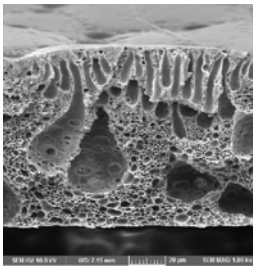
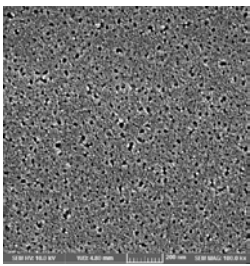
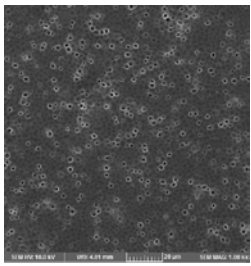
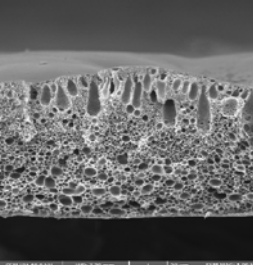
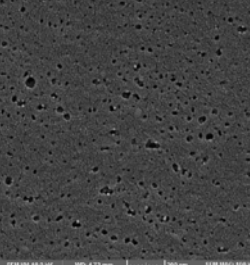
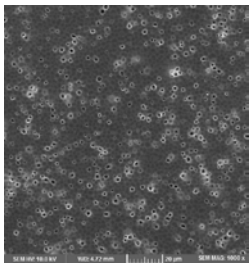
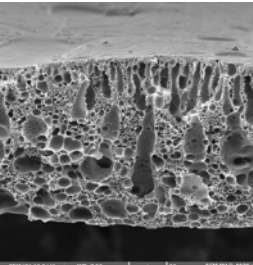
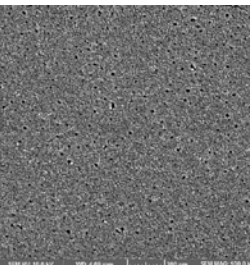
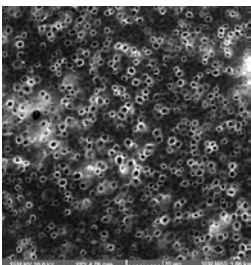
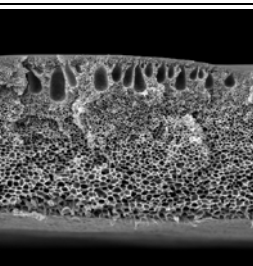
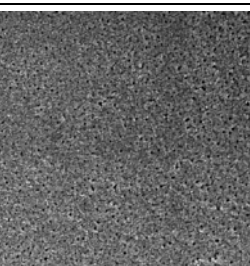
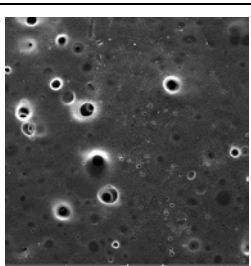
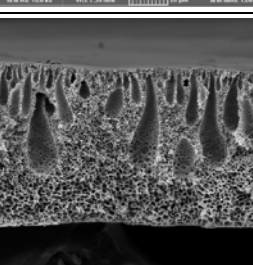
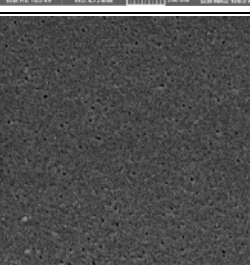
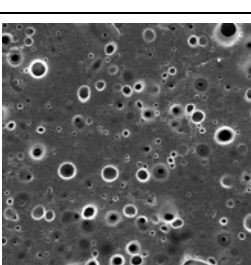
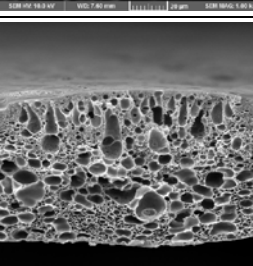
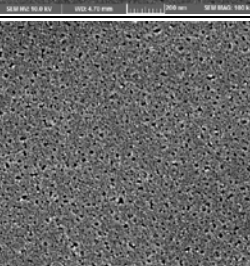
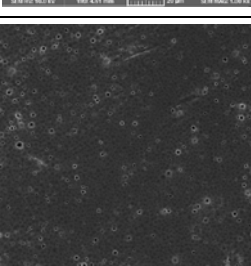
For porous membrane preparation, PVP, PEG and PEO are commonly used as relatively hydrophilic additives. Span 80 was used as a non-ionic surfactant with hydrophile-lipophile value (HLB) of 4.3 in polymer solutions [42]. It was claimed that the addition of PEGs commonly inhibits macrovoids formation, improves pores interconnectivity and increases the membrane porosity [43]. This was proved in this study as can be seen in Fig. 10 for the membrane PEG-0.2-A. The addition of PEG400 resulted in an effective transformation of finger-like and macrovoids structure to more sponge-like or honey-comb structure. According to the  $R_{HSP}$  value of PEG shown in Table 4, the affinity of PEG-CF ( $R_{HSP}=3.5$ ) is better than that of PEG-acetone or acetone/water mixture ( $R_{HSP}=10.65$  and  $18.02$ , respectively). The less exchange rate of PEG with the coagulants caused a slower outflow rate of PEG, slower inflow of non-solvent and therefore slower outflow of solvent during membrane formation inhibiting macrovoids formation as a consequence. As a result, the porosity was decreased (PEG-0.2-A membrane in Fig. 10 and Table 5). Compared to other membrane samples, using PEG400 as additive caused the reduction of the mean pore size (196 nm for PEG-0.2-A vs. 211 nm and 265 nm for CF-NA-A and PVP-0.2-AW, respectively, Table 5) and the enhancement of  $LEP$  (2.82 bar for PEG-0.2-A vs. 1.96 bar for CF-NA-A and 1.67



bar for PVP-0.2-AW, respectively, Table 5). It is worth noting that the precipitation kinetic and consequently the exchange rate of solvent and non-solvent were affected by the hydrophilic nature of PEG400. Kim and Lee [44] also reported that PEGs of small molecular weights, like PEG200 and PEG400, act as a pore-reducing agent in asymmetric polyetherimide (PEI) membranes.

As can be seen in Fig. 10, the finger-like structure of the CF-NA-A was transformed to the sponge-like structure up on the addition of PEO. Both PEO and PEG have the same constitutional units but with different synthesis methods, which results in a higher molecular weight for PEO. Therefore, compared to PEG400, PEO has a slower dissolution rate in a solvent due to its higher molecular weight resulting in a higher viscosity polymer solution. In this case, as it is mentioned earlier, because of the enhanced viscosity, the kinetic effect may become predominant over the thermodynamic effect and the solvent/non-solvent interdiffusion is delayed. It was proposed that under this delayed condition, the top layer has low porosity with low degree of pore interconnectivity [45]. In addition, the suppression of macrovoids in the sublayer could also be observed [46]. Thus, it may be expected a membrane with low porosity and permeability. As a result, the porosity of the membrane PEO-0.2-A was lower than that of PEG-0.2-A membrane (69.5% and 75.8%, respectively, Table 5). Gao et al. [47] also investigated the transition morphology of Poly (L-lactic acid) (PLLA) flat sheet membrane from finger-like pores to inter-connected pores by using PEO and PEG with different molecular weights. It was proved that the increase of the molecular weight of polymer chain (i.e. using PEO instead of PEG), changed the phase separation process to spinodal decomposition (SD) mechanism and the membrane structure transformed from finger-like to inter-connected network by increasing the concentration (wt%) of PEO in the casting solution. In addition, the  $R_{HSP}$  value of PEO-acetone is higher than that of PEG-acetone (i.e. less affinity of PEO towards the non-solvent compared to PEG400, Table 4). In this

1 regards, the less porosity and pore size of PEO-0.2-A compared to PEG-0.2-A is also justified.  
2 The  $R_{HSP}$  data related to PEO also showed less affinity of PEO with the non-solvents (both acetone  
3 and acetone/water) and consequently less diffusion rate of PEO to the coagulants. The less  
4 exchange rate of PEO with the coagulants caused a slower outflow rate of PEO and inflow rate of  
5 the non-solvent during membrane formation inhibiting therefore macrovoids formation. In  
6 contrast, the  $R_{HSP}$  data related with PVP showed more affinity of PVP with the non-solvents rather  
7 than with the solvent (Table 4). This resulted in an intense phase separation process between the  
8 polymer-rich phase and the polymer-lean phase and consequently macrovoids formation. In fact,  
9 the porosity and mean pore size of PVP-0.2-AW (79.4% and 265 nm, respectively, Table 5) are  
10 greater than those of PEO-0.2-A and PEO-0.2-AW (69.5%, 61.3% and 254, 241 nm, respectively,  
11 Table 5).

Membrane code	Cross section ( $\times 1,000$ magnification)	Top surface ( $\times 100,000$ magnification)	Bottom surface ( $\times 1,000$ magnification)
CF-NA-A			
PEG-0.2-A			
PEG-0.2-AW			
PEO-0.2-A			
PEO-0.2-AW			
Span 80-0.2-A			

1 **Fig. 10.** FESEM images of COP membranes prepared by different types of additives.

The additive Span 80 was selected to study COP phase separation because of its higher hydrophobic character compared to the other used additives. It is worth mentioning that the affinity of a given surfactant to water or oil can be expressed by the HLB value, The less the HLB value, the more hydrophobicity of the surfactant. Moreover, the Span 80 is a nonionic surfactant according to the form of its dissociation in the solvent. Since the addition of a surfactant affects the interfacial properties between coagulant and polymer solution, the membrane structure might be drastically changed. The membrane Span 80-0.2-A exhibits elongated macrovoids through its cross section and small finger like macrovoids near the top surface. Ge et al. [48] concluded that the addition of low HLB surfactant to the casting solution slowed down the rate of phase separation and increases its stability. It was also observed an increase of macrovoids in the sublayer of polysulfone membranes sublayer with the increase of HLB value of the surfactants. Tsai et al. [42] investigated the effect of a wide variety of surfactants with different HLB's and concluded that for a hydrophilic coagulant, hydrophilic surfactant enhanced macrovoids formation and a lipophilic surfactant was more effective to change the structure of membrane when the coagulant was lipophilic. In the present case, the affinity of Span 80-coagulant ( $R_{HSP}=5.5$ ) and solvent ( $R_{HSP}=6.6$ ) is almost similar according to the  $R_{HSP}$  values (Table 4). On the other hand, acetone is slightly lipophilic. As a result, the finger-like macrovoids of Span 80-0.2-A membrane was reduced and the spongy structure changed to big pores across the membrane. Consequently, the mean pore size of Span 80-0.2-A (239 nm) turned to be bigger than that of CF-NA-A, PEG-0.2-A and PEG-0.2-AW (211,196, 158 nm, respectively, Table 5) and smaller than that of PVP-0.2-AW, PEO-0.2-A and PEO-0.2-AW membranes (265, 254 and 241 nm, respectively, Table 5).

The addition of both PEO and Span 80 reduced the membrane porosity (69.5% and 50.0% for PEO-0.2-A and Span 80-0.2-A, respectively, vs. 80.3% for CF-NA-A) but increased the  $LEP$  value

(3.1 bar and 4.55 bar for PEO-0.2-A and Span 80-0.2-A, respectively, vs. 1.96 bar for CF-NA-A). This is attributed to the reduction of the bubble pore size. It was reported an enhancement of the membrane porous structure as the considered polymer solution became easier to phase separate [49]. Therefore, it is predictable the low porosity of the membranes PEO-0.2-A and Span 80-0.2-A because of the slower polymer phase separation compared to the membranes CF-NA-A and PVP-0.2-AW prepared with faster phase separation causing a finger-like structure.

As it is expected, the addition of a hydrophilic additive (PVP, PEG400 and PEO) may decrease the water contact angle of COP membranes if not leached out totally during membrane formation. The water contact angle was reduced in the order PEG-0.2-A > PEO-0.2-A > PVP-0.2-AW indicating that the PEG400 additive affects less the membrane hydrophobic character than the PVP additive. It must be pointed out that the addition of Span 80 also resulted in a decrease of the water contact angle even more than that of PEG400 (102.9° for PEG-0.2-A with 100° for PEO-0.2-A). Span 80 has both hydrophilic and lipophilic segments and the hydrophilic part may be responsible for the reduction of water contact angle. As it is mentioned before, the PEG400 additive is a small molecule so it can be easily removed by the non-solvent.

The addition of PEG400 increased the tensile strength, elongation at break and Young's modulus compared to CF-NA-A and PVP-0.2-Aw (Table 5) which may be due to the decrease of the porosity and change of the morphology of the CF-NA-A and PVP-0.2-AW membranes from finger-like structure to sponge-like structure with smaller macrovoids (see Table 5 and Fig. 10). As it was discussed previously, the addition of PEG400 suppressed the macrovoids and reduced the porosity (i.e. void volume fraction). Tsai et al. [50] also mentioned that cavities present in membranes could induce stress centralization forming weak points under an applied force and reduced mechanical properties.

1 According to the data shown in Table 5, the addition of PEO has a negative effect on the  
2 mechanical properties. Although the membrane PEO-0.2-A has a sponge-like structure with almost  
3 little macrovoids and less porosity, its tensile strength was decreased compared to the membrane  
4 CF-NA-A. The membrane Span 80-0.2-A exhibited low porosity compared to all other COP  
5 membrane but showed the highest mechanical properties (5.15 MPa for tensile strength, 7.94%  
6 elongation at break and 232 MPa for Young's modulus). This is attributed to its low porosity  
7 provided that a higher porosity weakens the membrane strength.

8 The short-term DCMD performance of the prepared porous COP membranes (CF-NA-A, PVP-  
9 0.2-AW, PEG-0.2-A, PEO-0.2-A and Span 80-0.2-A) was also summarized in Table 7. Based on  
10 the obtained results, the addition of PEG400 resulted in a slight decrease of the permeate flux  
11 compared to CF-NA-A and PVP-0.2-AW (Table 7) due to the decrease of its pore size and porosity  
12 (Table 5). Although the mean pore size of PEO-0.2-A is higher than that of PEG-0.2-A and CF-  
13 NA-A (254 nm vs. 196 nm and 211 nm, respectively, Table 5), an important decrease of the  
14 permeate flux was observed for PEO-0.2-A (see Table 7). As it is mentioned earlier, the addition  
15 of a high molecular weight PEO (i.e. membrane PEO-0.2-A) resulted in less porosity compared to  
16 PEG (i.e. membrane PEG-0.2-A) and a lower permeate flux. The lowest permeate flux was  
17 obtained for the membrane Span 80-0.2-A because of its lowest porosity (50%) compared to the  
18 membranes CF-NA-A, PVP-0.2-AW, PEG-0.2-A and even PEO-0.2-A. It must be pointed out that  
19 the salt separation factor of the membranes CF-NA-A, PVP-0.2-AW, PEG-0.2-A, PEO-0.2-A and  
20 Span 80-0.2-A was above 99.99% and the detected slight lower value 99.98% for the membrane  
21 PVP-0.3-AW was due to its lower *LEP* value (Table 5).

### 3.5.2. Effects of different coagulants on membrane characteristics and short-term DCMD performance

The effect of the coagulant on the morphological characteristics of two types of membranes (PEG-0.2-A/ PEG-0.2-AW and PEO-0.2-A/PEO-0.2-AW) was also investigated. With the addition of water (30 wt%) to acetone, the  $R_{HSP}$  value of the additive-mixed non-solvents increased in relation with that of the additive-acetone (see Table 4) indicating less affinity of the additive to the coagulant. As it is mentioned earlier, the slower outflow of the solvent caused suppression of macrovoids and consequently resulted in less porosity and smaller mean pore size. By comparing the membranes PEG-0.2-AW with PEG-0.2-A and PEO-0.2-AW with PEO-0.2-A, the addition of water to acetone resulted in lower porosity and pore size (i.e. 73%, 158 nm and 61.3%, 241nm for PEG-0.2-AW and PEO-0.2-AW, respectively, Table 5) than those of the membranes PEG-0.2-A and PEO-0.2-A, respectively (i.e. 75%, 196 nm and 69.5%, 254 nm, respectively, Table 5). Moreover, although the mean pore size of membranes prepared with acetone/water are less than that of the membranes prepared with acetone, a little increase of the bubble pore size was observed (Fig.10) and consequently resulted in a little decrease of  $LEP$  (Table 5).

Based on the obtained data of water contact angle of the membrane PEG-0.2-AW compared to PEG-0.2-A (103.8° and 102.9°, respectively, Table 5) and the membrane PEO-0.2-AW compared to PEO-0.2-A (99.7° and 98.8°, respectively, Table 5), the addition of water to acetone led to a slight increase of water contact angle may be due to the increase of the roughness parameters (Table 6). Furthermore, the addition of water to acetone increased the tensile strength, the elongation at break and Young's modulus of the membranes prepared with PEG as additive. In contrast, the mechanical properties of the membrane prepared with PEO were lower (i.e. the tensile strength of the membrane PEO-0.2-AW is 0.74 MPa) due to the fact that water intensified the

1 phase separation of the polymer solution inducing interfacial pores between the PEO and COP  
2 polymer.

3 The effect of the coagulant (acetone and acetone/water) on the DCMD performance was also  
4 shown in Table 7. Compared to the membranes prepared in acetone, bigger macrovoids appeared  
5 in the structure of the COP membranes prepared with PEG400 and PEO in the acetone-water  
6 mixture (i.e. a slight increase of the bubble pore size and decrease of the *LEP*). Based on the data  
7 shown in Table 7, the observed decrease of the permeate flux was due to the less porosity and pore  
8 size while the decrease of the salt separation factor was due to the slight reduction of the *LEP* of  
9 PEG, PEO-0.2-AW membranes compared to PEG, PEO-0.2-A membranes, respectively.

10 For sake of comparison, Table 8 lists the DCMD performance of some commercial and  
11 prepared membranes together with their characteristics. It can be seen that COP membranes exhibit  
12 DCMD permeate fluxes within those presented in Table 8. In fact, COP membranes showed  
13 slightly lower permeate fluxes than some commercial PVDF membranes (GVHP and HVHP) and  
14 PTFE membranes (TF200 and TF450), but higher than those of the commercial membranes (PE  
15 and MCE) and comparable to that of the commercially available PVDF membrane (PVDF22).



**Table 8.** Reported DCMD permeate flux ( $J$ ) of different types of commercial and prepared flat sheet membranes together with their characteristics ( $\mu_P$ : mean pore size;  $T_f$ : feed temperature;  $T_p$ : permeate temperature).

Membrane material	Thickness ( $\mu\text{m}$ )	$\mu_P$ ( $\mu\text{m}$ )	LEP (bar)	$\varepsilon$ (%)	$\theta$ ( $^\circ$ )	$J$ ( $\text{kg/m}^2\text{h}$ )	Test conditions	Ref.
GVHP <sup>a</sup>	118	0.22	2	70.1	110	32.4 25.2	distilled water as feed; $T_f=70^\circ\text{C}$ ; $T_p=20^\circ\text{C}$ distilled water as feed; $T_f=70^\circ\text{C}$ ; $T_p=20^\circ\text{C}$	[51] [52]
HVHP <sup>a</sup>	116	0.45	1.1	71.3	-	38.8	distilled water as feed; $T_f=70^\circ\text{C}$ ; $T_p=20^\circ\text{C}$	[51]
PVDF22 <sup>a</sup>	125	0.22	2.17	75	127.2	25	distilled water as feed; $T_f=80^\circ\text{C}$ ; $T_p=20^\circ\text{C}$	[53]
PE <sup>b</sup>	38	0.3	>3	88	118	~18	distilled water as feed; $T_f=70^\circ\text{C}$ ; $T_p=30^\circ\text{C}$ Feed flow rate; 74 L/h	[54]
TF200 <sup>c</sup>	0.178	0.2	0.282	80	-	~20 ~25	Olive mill wastewater as feed; $T_f=70^\circ\text{C}$ ; $T_p=20^\circ\text{C}$ distilled water as feed; $T_f=60^\circ\text{C}$ ; $T_p=20^\circ\text{C}$	[55]
TF450 <sup>c</sup>	0.178	0.45	0.138	80	-	~23 ~27	Olive mill wastewater as feed; $T_f=70^\circ\text{C}$ ; $T_p=20^\circ\text{C}$ distilled water as feed; $T_f=60^\circ\text{C}$ ; $T_p=20^\circ\text{C}$	[55]
MCE <sup>d</sup>	100	0.05	~24.1	78	120.9	~8	NaCl 1 M as feed; $T_f=60^\circ\text{C}$ ; $T_p=20^\circ\text{C}$	[56]
PVDF-HFP	142	0.075	-	46	-	4.41	NaCl 0.1M as feed; $T_f=65^\circ\text{C}$ ; $T_p=25^\circ\text{C}$ (99.95% salt retention)	[16]
Unsupported PVDF (additive: acetone+ $\text{H}_3\text{PO}_4$ )	141	0.13	-	60	84	8.97	NaCl 35g/L as feed; $T_f=50^\circ\text{C}$ ; $T_p=20^\circ\text{C}$ (more than 99% of salt retention)	[57]
Supported PVDF (additive: acetone+ $\text{H}_3\text{PO}_4$ )	213	0.22	-	55	83	12.45	NaCl 35g/L as feed; $T_f=50^\circ\text{C}$ ; $T_p=20^\circ\text{C}$ (more than 99% of salt retention)	[57]
PVDF <sup>e</sup>	74	0.17	-	44	71.7	~22	NaCl 35g/L as feed; $T_f=70^\circ\text{C}$ ; $T_p=20^\circ\text{C}$	[58]
Commercial PVDF	132	0.2	0.2	-	130	~17.5	Ratio of feed salt concentration to the initial concentration=1.9 $T_f=70^\circ\text{C}$ ; $T_p=20^\circ\text{C}$	[59]
Micropillared- (MP <sup>f</sup> )PVDF	263	0.1	0.2	-	155	~22		
CF-NA-A	85	0.211	1.79	80	111	20 <sup>g</sup> /15 <sup>h</sup>	<sup>g</sup> Distilled water as feed; $T_f=70^\circ\text{C}$ ; $T_p=25^\circ\text{C}$ <sup>h</sup> NaCl 30g/L as feed; $T_f=70^\circ\text{C}$ ; $T_p=25^\circ\text{C}$	this study
PVP-0.2-AW	70	0.265	1.67	79.4	93.5	16 <sup>g</sup> /13 <sup>h</sup>		
PEG-0.2-A	75	0.196	2.82	75.8	102.9	14 <sup>g</sup> /12.3 <sup>h</sup>		
PEG-0.2-AW	75	0.158	2.04	73.3	103.8	11 <sup>g</sup> /8 <sup>h</sup>		

<sup>a</sup> Commercial polyvinylidene fluoride (PVDF) flat-sheet membranes supplied by Millipore.

<sup>b</sup> Polyethylene (PE) based flat-sheet membrane supplied by Aquastill®, Netherlands.

<sup>c</sup> Polytetrafluoroethylene supported by a polypropylene net supplied by Gelman.

<sup>d</sup> Flat-sheet ultrafiltration mixed cellulose ester membrane supplied by Millipore.

<sup>e</sup> Membrane fabricated by novel method (bottom-up method) [58]

<sup>f</sup> Membrane fabricated by micromolding phase separation [59]

### 3.6. Long-term DCMD performance

To assess the long-term DCMD desalination stability, the prepared membranes with good performance (CF-NA-A, PVP-0.2-AW, PEG-0.2-A and PEG-0.2-AW) were selected and tested as stated earlier. The obtained DCMD permeate flux and salt separation factor of these membranes were plotted in Fig. 11. The permeate flux of the COP membrane prepared without any additive (CF-NA-A) was  $14.50 \pm 0.36 \text{ kg/m}^2\text{.h}$ , which was maintained stable during 9 h of experimental test. However, the salt separation factor started to decline after 5 h although its value was greater than 99%. In fact, this membrane exhibited a good porosity (80%) and *LEP* value (1.96 bar). After 9 h of DCMD operation, the permeate flux was increased gradually to 25, 36 and 44  $\text{kg/m}^2\text{.h}$  during 12 h (for simplicity not shown in Fig.11) and it was observed that this membrane started first to be broken from its edges. Although the tensile strength of this membrane was reasonably good (Table 5) compared with other membranes found in the literature, its elongation at break was low and the sudden variation of the flowrate of the feed solution in the long-term DCMD set-up caused an increase of the hydrostatic pressure at the inlet of the membrane module feed side and the subsequent membrane breakage from its edges. It is promising that the COP membranes having a little bit more flexibility were not broken in the long-term DCMD set-up. However, the more flexible COP membranes, obtained in this study, were made by hydrophilic additive, which increases the risk of membrane pore wetting especially if this hydrophilic additive is not totally leached out from the membrane structure.

The membrane PVP-0.2-AW showed a stable permeate flux of  $12.63 \pm 0.55 \text{ kg/m}^2\text{.h}$  during 14 h of DCMD operation and a salt separation factor greater than 99.6 % during 11 h of operation after which it started to decline reaching 99.2% after 14 h. In contrast to the membrane CF-NA-A,

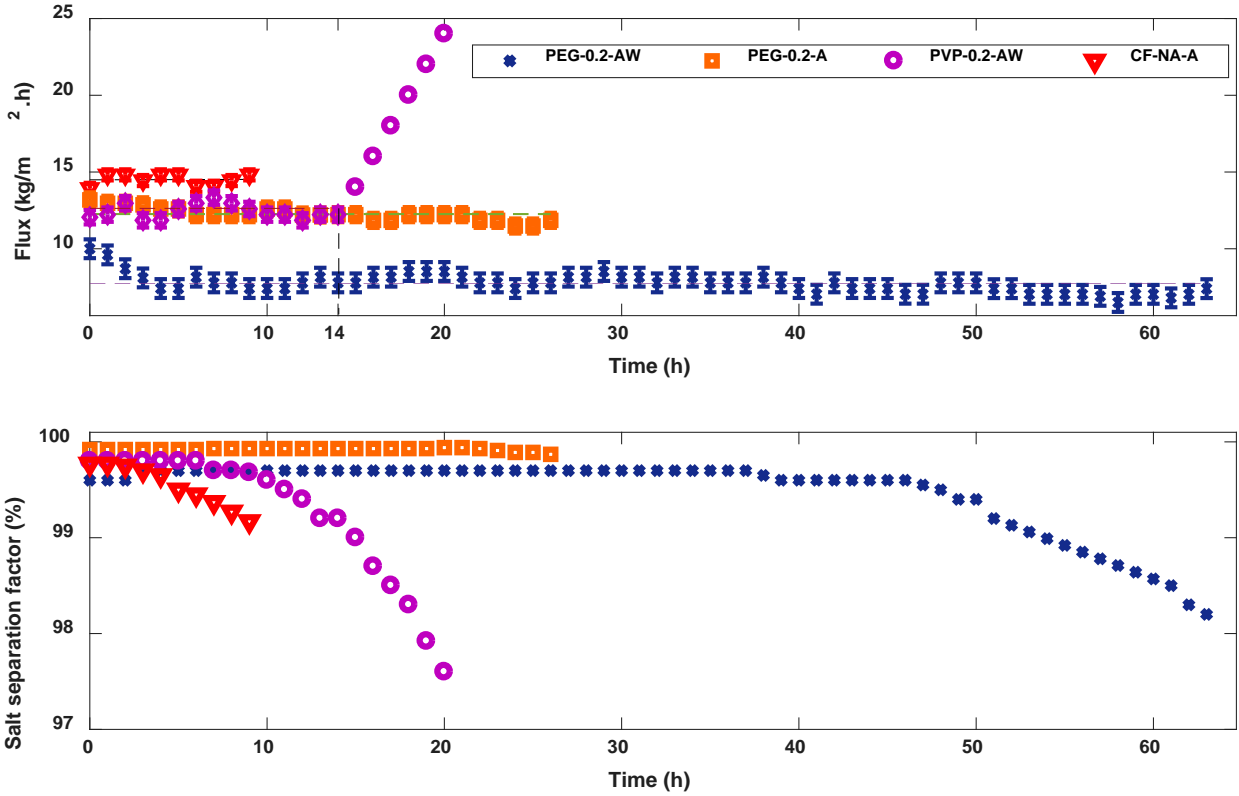
the membrane PVP-0.2-AW was not broken. This may be attributed to its better flexibility due to its better elongation at break. The Young's modulus ( $E$ ), which represents the membrane stiffness, was found to be lower, 190 MPa for the membrane PVP-0.2-AW compared to the other membranes (CF-NA-A, PEG-0.2-A and PEG-0.2-AW, Table 5). However, it was detected a gradual increase of the permeate flux of the membrane PVP-0.2-AW with time up to 24 kg/m<sup>2</sup>.h after 20 h DCMD operation and the salt separation factor was decreased to 97.6%. This indicated that pore wetting occurred, which is also confirmed by its low  $LEP$  value (i.e. 1.67 bar).

The permeate flux and the salt rejection of the membrane PEG-0.2-A exhibited good desalination performance during 26 h of the operation. The permeate flux was maintained stable at  $12.26 \pm 0.45$  kg/m<sup>2</sup>.h without any change of the salt separation factor (i.e. 99.9%) mainly due to the good  $LEP$  of this membrane (2.82 bar). However, after this time the permeate flux suddenly increased because the membrane was broken similar to what occurred to the membrane CF-NA-A although its mechanical properties are better than those of the membrane CF-NA-A. As it was shown in Table 5, the membrane PEG-0.2-A had almost similar tensile strength and Young's modulus to those of the membrane CF-NA-A and some improvement in the elongation at break.

Compared to the previous selected COP membranes, the lowest and longer stable DCMD permeate flux was obtained for the membrane PEG-0.2-AW ( $7.74 \pm 0.61$  kg/m<sup>2</sup>.h) with a salt separation factor greater than 99.4% during 50 h operating time. The low permeate flux of this membrane was expected due to its lower porosity (73.3%) and pore size (158 nm) compared to the other membranes, CF-NA-A, PVP-0.2-AW and PEG-0.2-A. After 55 h of operation, the salt separation factor decreased to 98.9% and at 63 h it reached 98.2%. This membrane was not broken as it showed the best mechanical properties compared to the membranes CF-NA-A, PVP-0.2-AW

1 and PEG-0.2-A (Table 5). Based on the low affinity of PEG-NS<sub>m</sub> ( $R_{HSP} = 18.02$ ) compared to  
2 PEG-S ( $R_{HSP} = 3.5$ ) and the higher  $R_{HSP}$  value of PEG-A/W compared to that of PEG-A (18.02  
3 *versus* 10.65), the acetone/water was a weak coagulant compared to acetone. It was reported that  
4 the use of a weak coagulant caused a slow solidification hindering therefore the formation of walls  
5 between small droplets [60]. Then, some small droplets were combined to form a large droplet  
6 resulting in large drop-shaped cavities in the sub-layer (Fig. 10) [60]. In this type of structure, the  
7 complete removal of additive may be more difficult. Therefore, the reduction of the salt separation  
8 factor may be attributed to membrane wetting due to the presence of some additives in the  
9 membrane structure.

10 It is promising that the elongation at break of the COP membranes should be improved further  
11 in order to prevent their breakage (compare the elongation at break of CF-NA-A with that of PVP-  
12 0.2-AW in Table 5). The addition of hydrophilic additives improves the flexibility of the  
13 membranes but it may result in membrane wetting if the additive is not fully leached out from the  
14 membrane matrix. The use of inorganic nanoparticles can be an effective method to improve both  
15 the elongation at break and MD performance and prevent wetting of COP membranes.



**Fig. 11.** Long-term DCMD performance of the COP membranes, CF-NA-A, PVP-0.2-AW, PEG-0.2-A and PEG-0.2-AW: permeate flux (a) and salt separation factor (b) as a function of operating time.

#### 4. Conclusions

Flat-sheet membranes were prepared for the first time from a novel and cheap engineering plastic (COP) via non-solvent induced phase separation technique (NIPS). Different types of solvents, coagulants and additives were examined and the resultant membranes were characterized for their use in desalination by direct contact membrane distillation (DCMD). It was found that the porosity of the prepared COP membrane lied in the range 50.0 – 82.3 %, their liquid entry pressure (*LEP*) values were in the range 1.2 – 4.5 bar, their hydrophobic character was confirmed with water contact angle values up to 111°, their mean pore size varied from 0.158  $\mu\text{m}$  to 0.265  $\mu\text{m}$  and their

tensile strength, elongation at break and Young's modulus were in the range 0.74-5.51 MPa, 3.3-7.9% and 29-237 MPa, respectively. These data are comparable with those of the porous membranes commonly used in MD technology.

Short-term DCMD tests of the COP membrane prepared without any additive showed the highest permeate flux (i.e. 20 kg.m<sup>-2</sup>.h<sup>-1</sup> and 15 kg.m<sup>-2</sup>.h<sup>-1</sup> when distilled water and 30 g/L sodium chloride aqueous solution were used as feed, respectively) with a high salt separation factor (i.e. 99.99%). However, this membrane didn't show a good long-term DCMD performance due to its high brittleness. This problem could be solved by using some additives, which resulted in a more flexible COP membrane. Long-term DCMD of some selected membranes prepared with PVP and PEG additives, showed pore wetting of the membrane prepared with PVP after 14 h of DCMD operation whereas the membrane prepared with PEG with different coagulants (acetone and acetone/water) exhibited more stable DCMD performance (i.e. 12.26 and 7.74 kg/m<sup>2</sup>.h when using acetone and acetone/water coagulants, respectively) during 24-50 h of desalination operation without any significant changes of the salt separation factor. In general, this research study proved that COP can be successfully used as a new candidate polymer in MD membrane engineering and MD applications. More research studies are needed to improve the MD performance of the new candidate COP membrane.

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